

Acid Rules:

Hydrogen and one other element:

ic acid – formed from the "ide" ion and hydrogen hydrosulfuric acid H₂S hydrochloric acid HCl

Hydrogen and two or more elements

ic acid – formed from the "ate" ion and hydrogen sulfuric acid H₂SO₄ Chloric acid HClO₃

ous acid – formed from the "ite" ion and hydrogen Chlorous acid sulfurous acid H₂SO₃

Prefixes – COVALENT ONLY!

1 – mono	2 – di	3 – tri	4 – tetra	5 – penta
6 – hexa	7 – hepta	8 – octa	9 – nona	10 – deca

Hydrates:

Hydrates – have water attached to them i.e. CuSO₄•5H₂O Name the ionic compound – copper(II)sulfate Then add the number of "hydrates" so pentahydrate Name - copper(II)sulfate pentahydrate

Remembering your Polyatomic Ions:

Memorize the "ate" ion

"-ate" + 1 Oxygen	"perate"	SO ₅ ²⁻	persulfate
"-ate"	"-ate"	SO_4^{2-}	sulfate
"-ate" – 1 Oxygen	"-ite"	SO_3^{2-}	sulfite
"-ate" -2 Oxygen	"hypoite"	SO_2^{2-}	hyposulfite